

Practice:

Classify the compounds as ionic, non-polar covalent, polar covalent:

a. Cs-F

b. N-S

c. C-F

Jan 14-7:34 AM

Practice:

accel chem

Determine if the following describes an ionic or covalent compound :

1. Can be nonpolar due to equal electron sharing.
2. Formula must have balanced charges.
3. Formed by oppositely charged ions attracting.
4. Have very high melting points and boiling points.
5. Formed between a metal and a nonmetal.
6. Has bonding and unshared electron pairs.
7. Formed by a shared pair of electrons.

Jan 14-7:34 AM

Practice:

accel chem

Draw the Lewis structure, identify the shape and the molecular polarity:

a. HCN

b. NF_3

Jan 14-7:34 AM

Practice:

Name the following compounds:

a. MnCO_3

b. P_2O_5

c. SrF_2

Jan 14-7:34 AM

Practice:

Write the formula for the following compounds:

- nitrogen monoxide
- magnesium hydroxide
- tricarbon octahydride

Jan 14-7:34 AM

Practice:

accel chem

Determine the main intermolecular force acting on the following:

- NH_3
- F_2
- HCN

Jan 14-7:34 AM

Practice:

accel chem

Put the following in order from least to highest boiling point: N_2 , CaCl_2 , H_2O

Jan 14-7:34 AM

Practice:

Describe the differences in bond strength between ionic and covalent compounds. Provide reasons as to why the strength differs.

Jan 14-7:34 AM