Practice:

Neon has three isotopes. Neon-20 has a mass of 19.992 amu and abundance of 90.48%, Neon-21 has a mass of 20.994amu and an abundance of 0.27%, and Neon-22 has a mass of 21.991 amu and an abundance of 9.25%. What is the average atomic mass for neon? Show formula, set-up, and answer with units.

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Practice:

Copper used in electric wires comes in two flavors (isotopes): ^{63}Cu and ^{65}Cu. The average atomic mass between these two isotopes is 63.546 amu. Calculate the percent abundance of each isotope.

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Practice:

Give the number of protons, neutrons and electrons in the following:

Isotope	Protons	Neutrons	Electrons
uranium-235			
$^{59}_{28}Ni$			

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Practice:

How would you represent an isotope of carbon with seven neutrons?

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Practice:

Write the complete equation for the alpha fusion of radon – 198.

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Practice:

Arrange the following cations (positive ions) in order of radius: Rb^{+}, Be^{2+}, and Sr^{2+}. increasing ionic

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Practice:

Arrange the following elements in each set in order of decreasing ionization energy: Si, Cl, F

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Practice:

Arrange the following elements in order of decreasing atomic radius: O, N, F

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Practice:

Determine which element has the following electron configuration:

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^3$

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Practice:

What is the noble-gas configuration for silicon?

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Practice:

Draw the electron-dot structure for arsenic.

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Practice:

Write the electron configuration for an ion of Fluorine (F^-).

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Practice:

Draw the orbital diagram for P.

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