

Review

What solute will lower the freezing point of a solvent the most? NaCl, CH₄, Mg(OH)₂

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Review

What substances are likely to dissolve in water?

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Review

Describe what happens when you add 1 crystal to the following solutions:

saturated solution:

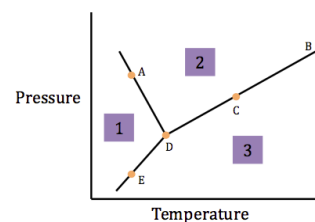
supersaturated solution:

unsaturated solution:

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Review

- Determine the phase for each of the numbers labeled in the diagram below.
- What phase change is happening at letter c?
- What term is used to describe point D?
- What change is happening to the solution when it moves from 1 to 3?



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Review

What will the freezing point be of a solution containing 40 grams of NaOH in 500 kg of water? ($k_f = 1.86^\circ\text{C}/\text{m}$)

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What is the molality of a solution that has 50 moles of solute in 1000 mL of water?

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Review

What is the volume of a 0.5 M solution that has 10 moles of solute?

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