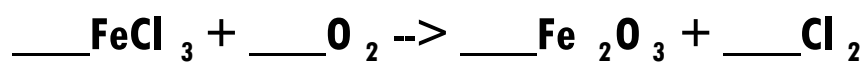


Practice:

How many grams of oxygen gas is required to react with 10.0 moles of iron (III) chloride?



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Practice:

Determine the number of significant figures in the measurements below.

- a. 0.0000236
- b. 1000
- c. 6300.
- d. 6020
- e. 10.021

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Practice:

Calculate the following and put into correct significant figures.

a. $12.53 - 6.1 + 2.01$

b. $0.206 + 0.002$

c. $16.200/2.00000$

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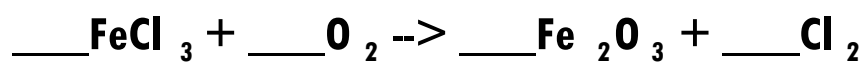
Practice:

What is the molar mass of NaCl?

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Practice:

How many moles of chlorine gas can be produced if 4.0 moles of FeCl_3 react with 4.0 moles of O_2 ?



- What is the limiting reactant?
- How much excess do you have?

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Practice:

accel chem

The percent composition of a compound was found to be 63.5% silver, 8.2% nitrogen and 28.3% oxygen. Determine the compound's empirical formula.

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Practice:

accel chem

The empirical formula for the trichloroisocyanuric acid, the active ingredient in many household bleaches, is OCNCl . The molar mass of this compound is 232.41 g/mol. What is the molecular formula of trichloroisocyanuric acid?

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Practice:

accel chem

Determine the percent yield of the experiment below.



Initial Mass of $\text{Al}(\text{OH})_3$	1.00 grams
Mass of Beaker and Reactants	15.650 grams
Mass of Solid Product ($\text{Al}_2(\text{SO}_4)_3$)	15.000 grams
Mass of gas product (H_2O)	

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