Unit 5 In-Class Review Name: _____ Period: _____

1. Silver nitrate (aq) + copper (s) → copper (II) nitrate (aq) + silver (s)

a. What type of reaction is occurring? How do you know?

b. What do you have to look at to know if this reaction will occur or not?

c. What is the precipitate formed in the reaction?

2. Propane (C_3H_8) burns in the presence of oxygen gas →

a. What type of reaction is occurring? How do you know?

b. What is the formula for oxygen in this reaction? Why?

c. What is the coefficient for carbon dioxide in the balance chemical equation?

3. Iron (III) chloride (aq) + sodium hydroxide (aq) → net ionic equation

a. What type of reaction is occurring? How do you know?

b. What type of compounds are the reactants and products? How do you know?

c. What does the "aq" stand for in the reaction? Why is knowing the states of matter important?

4. Hydrogen (g) + nitrogen (g) → ammonia (g)

a. What type of reaction is occurring? How do you know?

b. What product(s) is/are formed?

c. Are there any subscripts used in this equation? What do they tell us?

5. Magnesium oxide (s) is heated and breaks down into its elements.

a. What type of reaction is occurring? How do you know?

b. What is/are the reactant(s) in this reaction?

c. Why do chemical equations have to be balanced?

6. Write the molecular, complete, & net ionic equation when solutions of silver sulfate and magnesium chloride react.

- What is the precipitate in this reaction?