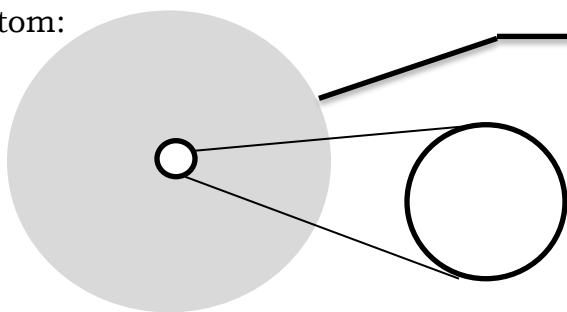


Subatomic Particles Guided Notes

Name: _____ Period: ____

Parts of the atom:



-Empty space around nucleus
 -Where electron are located
 -Majority of the _____ of atom

_____ -Center of the atom
 - Contains the protons & neutrons
 -Majority of the _____ of atom

Sub-Atomic Particles:

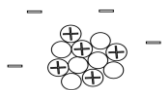
Name			
Symbol			
Charge			
Mass**			
Location			
Purpose			

** AMU stands for _____, which is:

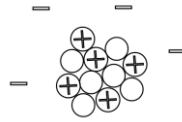
Important Terms:

Atomic Number	The number of _____ in an atom Like the ID # because it determines the _____ of an element.
Mass Number	The mass of 1 atom of an element Mass number = the # of _____ + the # of _____
Neutral	No Charge In a _____ the # of protons = the # of electrons

1. Which particle can tell you the identity of an atom? A. proton B. Neutron C. Electron



2. What element is this?



3. What is the mass number of this atom?
4. If an atom of carbon has a mass number of 13, how many neutrons does it have?
5. How many electrons are in a neutral atom of magnesium?

Homework: Complete the chart

Element:	Protons	Neutrons	Electrons	Atomic Number	Mass Number
Potassium	19	20			
	14	16			30
Cobalt		31			
	7				15