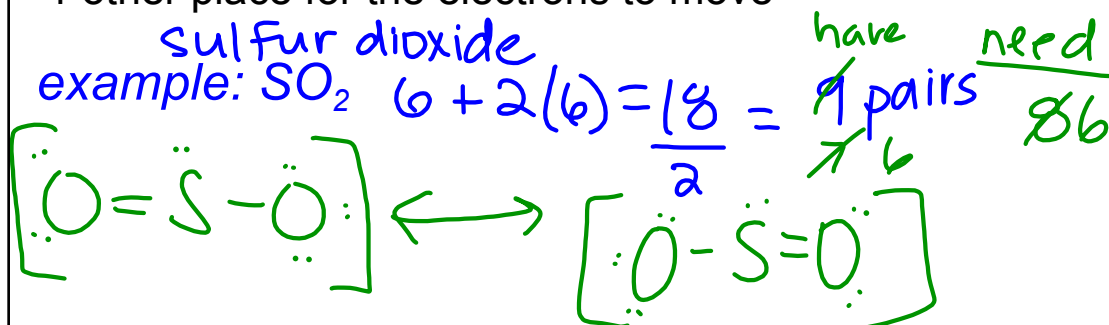


Resonance Structures:

-- have the same arrangement of elements, but a different arrangement of electrons

— shared pair .. lone pairs

-- need to have at least one multiple bond and at least 1 other place for the electrons to move



Oct 17-9:00 AM

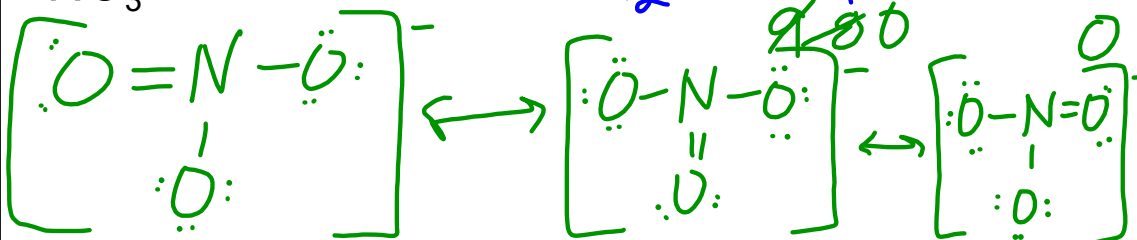
Polyatomic Ion Lewis Structures:

- same as drawing other Lewis Structures
- negative ions, add electrons
- positive ions, subtract electrons
- bracket and put the charge in the upper right
- may also have resonance

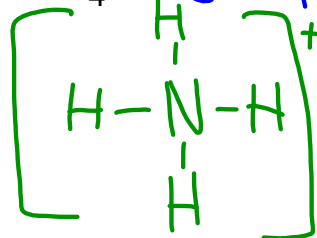
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Examples:

nitrate
 NO_3^- $5 + 3(6) + 1 = \frac{24}{2} = 12$ pairs ~~10~~ ~~8~~ ~~6~~



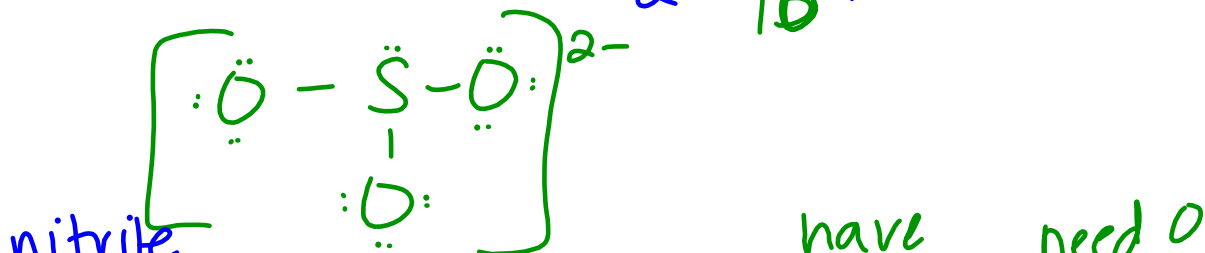
NH_4^+ $5 + 4(1) - 1 = \frac{8}{2} = 4$ pairs



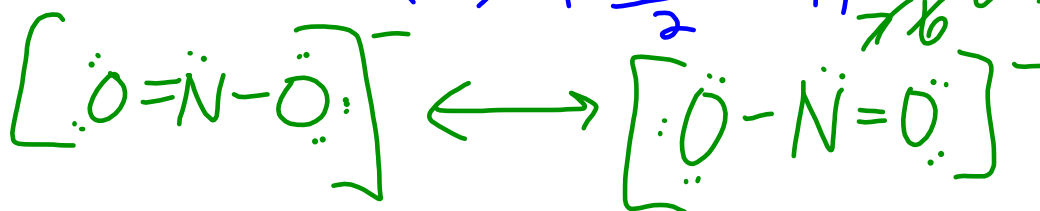
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Practice:

SO_3^{2-} $6 + 3(6) + 2 = \frac{26}{2} = 13$ pairs ~~10~~ ~~8~~ ~~6~~



nitrite
 NO_2^- $5 + 2(6) + 1 = \frac{18}{2} = 9$ pairs ~~10~~ ~~8~~ ~~6~~



Oct 17-9:10 AM