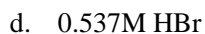
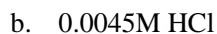
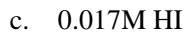
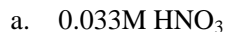
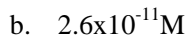
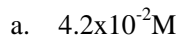
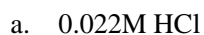
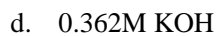
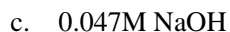
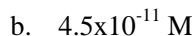
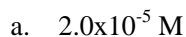


pH and pOH Practice (Accelerated Chemistry)

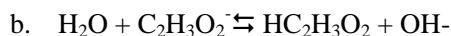
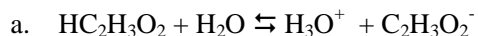
Name: _____ Pd: _____

Show formula, setup and answer with units if appropriate.

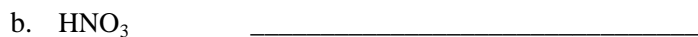
1. Determine the pH of the following acid solutions:

2. What is the pH of a solution if its [H⁺] is:3. Calculate the [OH⁻] for the following acids:4. Determine the pH if the [OH⁻] is:

8. Identify each as an acid, base, conjugate acid and conjugate base. You may use BA, BB, ca, cb.



9. Classify each of these as an Arrhenius acid or base:

10. What is true about the relative concentrations of hydrogen ions [H⁺] & hydroxide ions [OH⁻] in each of these solutions:

a. Basic _____

b. Acidic _____

c. Neutral _____