Percent Composition Practice

Name/Period: _____

Guided Notes:

1. Define percent composition.

2. What is the general equation used to solve for percent composition?

3. What do the percentages of the elements in a compound need to add up to?

4. Find the percent by mass of *each* element in $NaNO_3$ using the equation in #2. **SHOW YOUR WORK!!!**

Complete the following percent composition practice problems. Show *all* of your work!

1. Determine the percent composition of calcium in calcium chloride ($CaCl_2$).

2. Calculate the percent composition of sodium sulfate (Na_2SO_4).

3. Which has the larger percent by mass of sulfur, H_2SO_3 or $H_2S_2O_8$?

4. Which has the higher percent hydrogen- phosphoric acid (H_3PO_4) or hydrochloric acid (HCl)?

Mole Conversion Practice:

5. How many grams are in 8.23 moles of sodium carbonate?

6. How many grams are in 7.23×10^{24} molecules of carbon dioxide?

7. How many moles are in 45.8 grams of copper?

8. How many moles are in 7.8×10^{23} ions of Ca^{2+}?

9. How many grams are in 2.3×10^{23} formula units of $Al_2(CO_3)_3$?

10. How many ions are in 50.1 grams of NH_4^+?