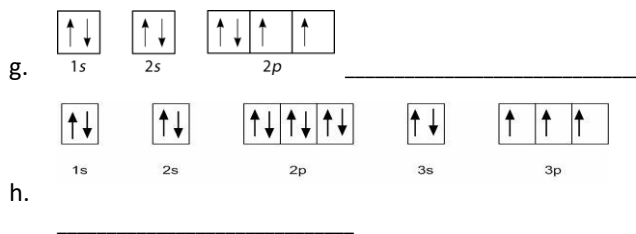


# Mixed Configurations Practice

Name: \_\_\_\_\_ Period: \_\_\_\_\_

1. Identify the following elements based on the information given:

- a.  $3s^2$  \_\_\_\_\_
- b.  $6p^3$  \_\_\_\_\_
- c.  $[\text{Xe}]6s^24f^{14}5d^4$  \_\_\_\_\_
- d.  $[\text{Ar}]4s^23d^3$  \_\_\_\_\_
- e.  $1s^22s^22p^3$  \_\_\_\_\_
- f.  $1s^22s^22p^63s^23p^64s^23d^7$  \_\_\_\_\_



2. Write the ground state electron configurations for the following elements:

- a. Gallium \_\_\_\_\_
- b. Osmium \_\_\_\_\_
- c. Bohrium \_\_\_\_\_
- d. Arsenic \_\_\_\_\_

3. Draw the orbital diagram configuration for the following elements:

- a. Potassium
- b. Sulfur
- c. Zinc
- d. Aluminum

4. Write the ending configuration for the following elements:

- a. Scandium \_\_\_\_\_
- b. Astatine \_\_\_\_\_
- c. Krypton \_\_\_\_\_
- d. Calcium \_\_\_\_\_
- e. Francium \_\_\_\_\_
- f. Iridium \_\_\_\_\_

5. Write the electron configurations for the following:

- a. Na \_\_\_\_\_
- b. S \_\_\_\_\_
- c. B \_\_\_\_\_
- d. Co \_\_\_\_\_