

Guided Notes: K_{sp}

Name: _____ Pd.: _____

Solubility Product Constant – K_{sp}

- K_{sp}:
 - General Equation:
 - Since the reactant is ALWAYS a _____, K_{sp} = _____
 - b and c are the _____ on the ions
 - The **smaller** K_{sp} is the _____ soluble salt
 - K_{sp} can be used to calculate the _____ of _____.

Practice - K_{sp}

1. Write the K_{sp} expression for the solvation of Ag₂SO₄.

First, determine the ions that will be formed:

Put the ions in the K_{sp} expression (must include charges!):

Use the coefficients to determine how many moles of each ion will be formed. Put those numbers in for b & c (as exponents):

(if the exponent is _____, it is not used in the expression)

2. Write the K_{sp} expression for the solvation of magnesium hydroxide. Formula: _____

3. Write the K_{sp} expression for the solvation of calcium phosphate. Formula: _____