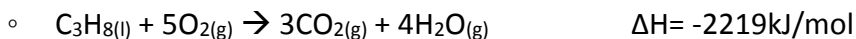
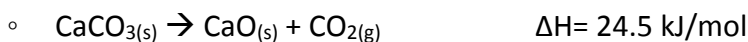


**Guided Notes: Enthalpy and Thermochemical Equations****Enthalpy**

- heat energy (\_\_\_\_) measured in Joules (\_\_\_\_) or kilojoules (\_\_\_\_)
- \_\_\_\_\_ = endothermic
- \_\_\_\_\_ = exothermic

**Thermochemical Equations:**

A balanced chemical equation that shows the \_\_\_\_\_ (enthalpy) either \_\_\_\_\_ or \_\_\_\_\_ in the reaction

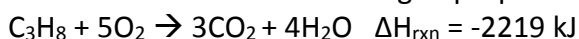


\_\_\_\_\_ Reactions:  $+\Delta H_{\text{rxn}}$  (\_\_\_\_\_ energy)

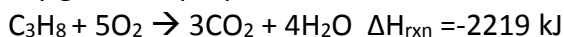
\_\_\_\_\_ Reactions:  $-\Delta H_{\text{rxn}}$  (\_\_\_\_\_ energy)

**Practice:**

- How much heat is released when 5.7 g of propane are burned?



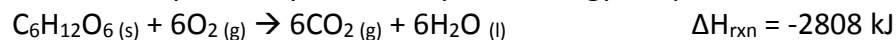
- How many grams of propane would be needed to produce 255kJ of energy?



- How many grams of carbon dioxide are made when we produce 457 kJ?

**Check For Understanding**

- When glucose is burned in your body, how many kJ of energy are produced for every 1.0g of glucose?



- Is this reaction endothermic or exothermic? \_\_\_\_\_