

Guided Note: Empirical and Molecular Formulas

Empirical Formulas

Empirical Formulas: _____

Steps for Finding Empirical Formulas:

1. The percent is in 100 g so, _____.
2. Calculate the number of _____.
3. Divide by the _____.
 - a. Number within 0.1 _____.
 - b. If you don't get a whole number, _____.

Ex:

4. The whole numbers become _____.

Practice:

1. The percent composition of an unknown compound is found to be 34.53% Zn, 14.79% N and 50.69% O. Determine the compounds empirical formula. (Show all work in your notes!)

2. Calculate the empirical formula for a compound whose analysis is 74.97% Al and 25.03% C. (Show all work in your notes!)

Molecular Formulas

Molecular Formulas: _____

2 Things you need to know to determine the molecular formula:

1. _____
2. _____

Equation to find molecular formula:

