

# Net Ionic Equations

Name/Per: Key

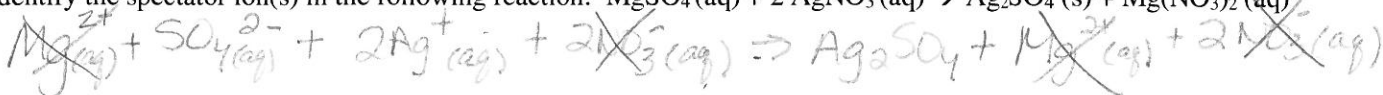
1. Explain why the term spectator ion is used.

*not actively participating in the reaction*

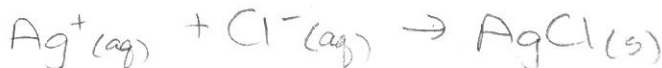
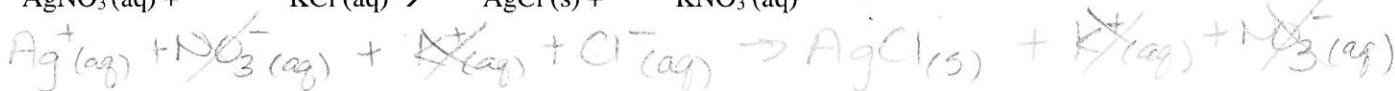
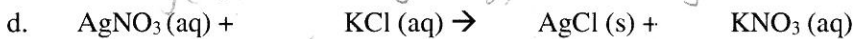
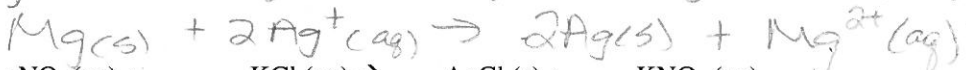
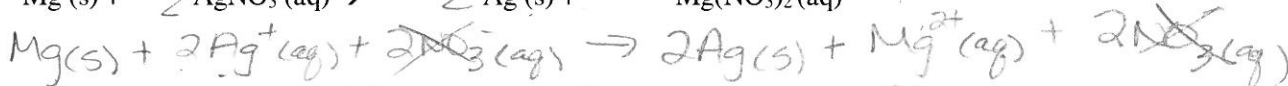
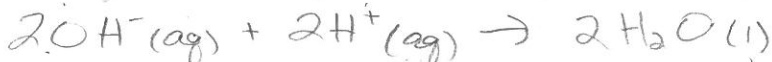
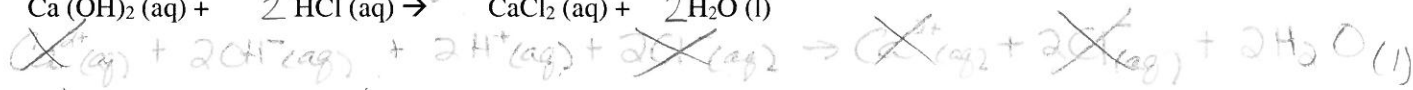
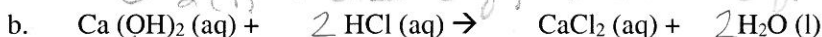
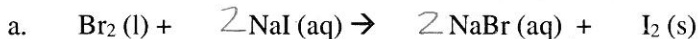
2. What chemicals are present in a net ionic equation?

*only the ions and substances that actively participate in the reaction*

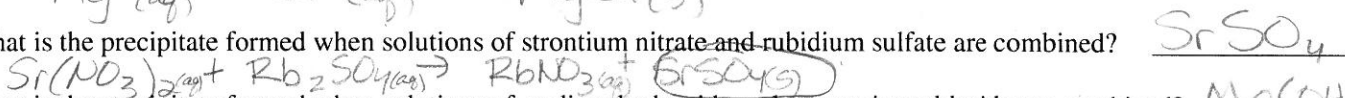
3. Identify the spectator ion(s) in the following reaction:  $MgSO_4(aq) + 2 AgNO_3(aq) \rightarrow Ag_2SO_4(s) + Mg(NO_3)_2(aq)$



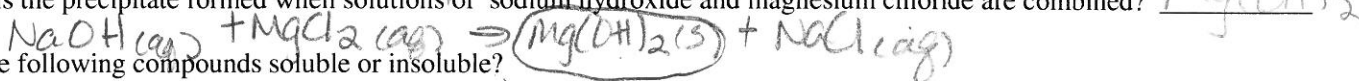
4. Balance the following molecular equations. Then write the complete and net ionic equation for each (show all steps):



5. What is the precipitate formed when solutions of strontium nitrate and rubidium sulfate are combined?



6. What is the precipitate formed when solutions of sodium hydroxide and magnesium chloride are combined?



7. Are the following compounds soluble or insoluble?

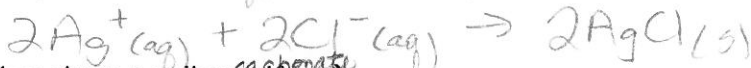
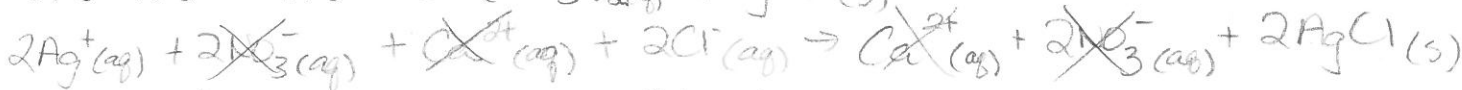
a. Calcium carbonate insoluble

c. Na<sub>2</sub>S soluble

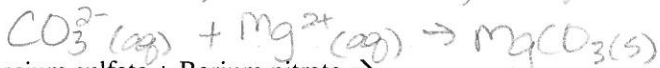
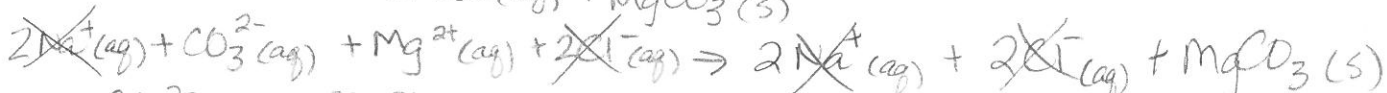
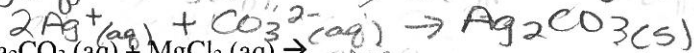
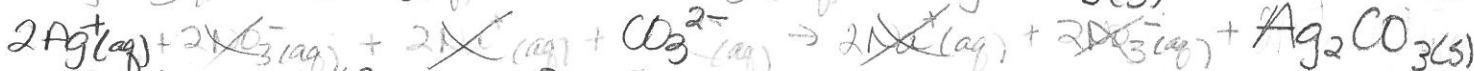
b. PbSO<sub>4</sub> insoluble

d. K<sub>2</sub>CO<sub>3</sub> soluble

8. Predict the products for each of the following reactions. Write the molecular, complete, and net ionic equations for each.



b. Silver nitrate + sodium carbonate  $\rightarrow$



d. Potassium sulfate + Barium nitrate  $\rightarrow$

