DOUBLE REPLACEMENT ACTIVITY-Accel

Name: _____Pd:___

For questions 1 and 2 balance the equation

1. Na_2CO_3 + HCl → NaCl + H_2CO_3

For question2, predict the 2nd product, and then balance the equation.

2. H_2SO_4 + $Ba(OH)_2$ → $BaSO_4$ + _____

For question 3, predict both products, and then balance the equation.

3. $CaCl_2$ + H_2O → _____ + _____

Write the balanced equation for the reaction between solutions of silver nitrate and copper(II) sulfate. Then balance the equation.

4. _____ + _____ → _____ + _____

Summing Up: Complete the following and balance the equation:

1. _____ $CaCO_3$ + ____ HCl → _____ + _____

2. _____ $Ba(C_2H_3O_2)_2$ + ____K_2SO_4 → _____+ _____

Teacher Initials:

_____

Reaction Types 2 Worksheet

Single Replacement:

1. Predict if the following reactions will occur (**Hint:** Use your activity series!).
   - If the reaction occurs, predict the products of the reaction and write the balanced chemical equation.
   - If no reaction occurs, write "No Reaction" after the arrow.

   a. _____ Mg + _____ $Pb(NO_3)_2$ →

   b. _____ Ni + _____ $Al_2(SO_4)_3$ →

2. Write a balanced equation for the reaction of solid copper (II) wire with a solution of silver nitrate.

Double Replacement:

3. Predict the products of the following reactions and balance the equations:

   a. _____ NaOH + _____H_2SO_4 →

   b. _____ $Sr(NO_3)_2$ + _____ Na_2SO_3 →

4. Write a balanced reaction for the reaction between solutions of potassium iodide and lead(II) nitrate.

5. Write a balanced reaction for the reaction between solutions of magnesium hydroxide and hydrochloric acid (HCl).

Synthesis and Decomposition:

6. Write a balanced equation for the reaction between carbon and sulfur (S_8) to form carbon disulfide.

7. Write a balanced equation for the synthesis reaction between lithium metal and liquid bromine.

8. Write a balanced equation for the decomposition of water into its elements.

Combustion:

9. Write a balanced equation for the combustion of ethanol (C_2H_5OH) in air.

10. Write a balanced equation for the reaction of butane (C_4H_{10}) with pure oxygen.

Balance each reaction and state the reaction type(s).

1. $NaOH + KNO_3 \rightarrow NaNO_3 + KOH$ _____

2. $Fe + NaBr \rightarrow FeBr_3 + Na$ _____

3. $CaSO_4 + Mg(OH)_2 \rightarrow Ca(OH)_2 + MgSO_4$ _____

4. $NH_4OH + HBr \rightarrow H_2O + NH_4Br$ _____

5. $Pb + O_2 \rightarrow PbO_2$ _____

6. $Na_2CO_3 \rightarrow Na_2O + CO_2$ _____

7. $Cr(s) + SnCl_4(aq) \rightarrow Sn(s) + CrCl_2(aq)$ _____

8. $RbCl(aq) + MgSO_4(aq) \rightarrow Rb_2SO_4(aq) + MgCl_2(aq)$ _____