

DOUBLE REPLACEMENT ACTIVITY-Accel

Name: _____ Pd: _____

For questions 1 and 2 balance the equation



For question 2, predict the 2nd product, and then balance the equation.



For question 3, predict both products, and then balance the equation.



Write the balanced equation for the reaction between solutions of silver nitrate and copper(II) sulfate. Then balance the equation.



Summing Up: Complete the following and balance the equation:



Teacher Initials:

Reaction Types 2 Worksheet

Single Replacement:

1. Predict if the following reactions will occur (**Hint:** Use your activity series!).

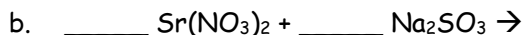
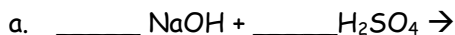
- If the reaction occurs, predict the products of the reaction and write the balanced chemical equation.
- If no reaction occurs, write "No Reaction" after the arrow.



2. Write a balanced equation for the reaction of solid copper (II) wire with a solution of silver nitrate.

Double Replacement:

3. Predict the products of the following reactions and balance the equations:



4. Write a balanced reaction for the reaction between solutions of potassium iodide and lead(II) nitrate.

5. Write a balanced reaction for the reaction between solutions of magnesium hydroxide and hydrochloric acid (HCl).

Synthesis and Decomposition:

6. Write a balanced equation for the reaction between carbon and sulfur (S_8) to form carbon disulfide.
7. Write a balanced equation for the synthesis reaction between lithium metal and liquid bromine.
8. Write a balanced equation for the decomposition of water into its elements.

Combustion:

9. Write a balanced equation for the combustion of ethanol (C_2H_5OH) in air.
10. Write a balanced equation for the reaction of butane (C_4H_{10}) with pure oxygen.

Balance each reaction and state the reaction type(s).

1. $NaOH + KNO_3 \rightarrow NaNO_3 + KOH$ _____
2. $Fe + NaBr \rightarrow FeBr_3 + Na$ _____
3. $CaSO_4 + Mg(OH)_2 \rightarrow Ca(OH)_2 + MgSO_4$ _____
4. $NH_4OH + HBr \rightarrow H_2O + NH_4Br$ _____
5. $Pb + O_2 \rightarrow PbO_2$ _____
6. $Na_2CO_3 \rightarrow Na_2O + CO_2$ _____
7. $Cr(s) + SnCl_4(aq) \rightarrow Sn(s) + CrCl_2(aq)$ _____
8. $RbCl(aq) + MgSO_4(aq) \rightarrow Rb_2SO_4(aq) + MgCl_2(aq)$ _____