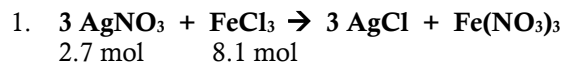


Limiting Reactant Practice- Accel.

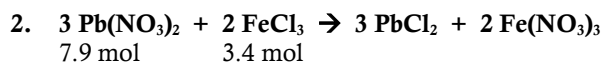
Name: _____ Pd: _____

For each of the following equations, show all of your work to determine the following:

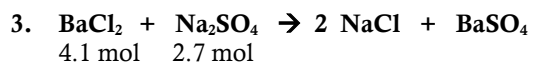
- a. Limiting reagent (reactant) b. Moles of
- specified*
- product produced c. Moles of excess reagent remaining



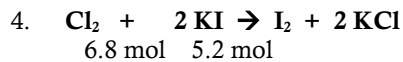
- a. LR = _____ b. _____ mol AgCl c. _____



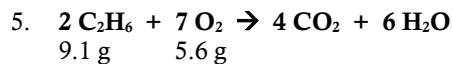
- a. LR = _____ b. _____ mol Fe(NO
- ₃
-)
- ₃
- c. _____



- a. LR = _____ b. _____ mol NaCl c. _____



- a. LR = _____ b. _____ mol KCl c. _____

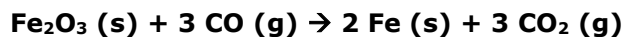


- a. LR = _____ b. _____ gCO
- ₂
- c. _____ g excess

Percent Yield Practice

For each problem below:

- balance the equation if needed
 - Show *all* of your work with units
 - calculate both theoretical yield & percent yield
 - Circle or box your answers with proper units
1. 84.8 g of iron (III) oxide reacts with an excess of carbon monoxide and 57.8 g of iron is produced.



2. 5.00 g of potassium chlorate is decomposed and 1.25 g of oxygen is produced.



3. 24.8 g of calcium carbonate can be decomposed, by heating, to produce 13.1 g of CaO.



4. In preparing a paint pigment of chrome yellow, PbCrO_4 , a student used 6.94 g of $\text{Pb}(\text{NO}_3)_2$. His actual yield of PbCrO_4 was 6.37 g.

