

Unit 9 Review Quiz: Accel Chem

- Which of the following is NOT a solution?
 - air
 - brass
 - orange juice with pulp**
 - lemonade
- A solution is made by dissolving salt in a beaker of water. The salt is referred to as the...
 - solution
 - solute**
 - solvent
 - dissolver
- When KCl is dissolved in water, the following will be produced:
 - $K + Cl$
 - $K + Cl_2$
 - $K^+ + Cl_2^-$
 - $K^+ + Cl^-$**
- What is the mass percent of an aqueous solution containing 15 g of KCl in 500 mL of water?
 - 0.7%
 - 3%**
 - 0.04%
 - 40.5%
- What is the molarity of solution that contains 1.5 moles in 2.5 L of solution?
 - 3.75 M
 - 0.6 m
 - 1.67 M
 - 0.6 M**
- A solution containing the maximum amount of dissolved solute is...
 - unsaturated
 - saturated**
 - super saturated
 - none of the above
- A student adds one crystal to a solution and observes additional crystals forming. This solution is...
 - unsaturated
 - saturated
 - supersaturated**
 - a chemical reaction
- Which of the following will increase the solubility of a gas in water?
 - increasing the temperature
 - increasing the pressure**
 - increasing the volume
 - decreasing the pressure
- How many moles of NaCl will be needed to make 500 mL of a 2.0 M aqueous solution?
 - 4 mol
 - 1.0 mol**
 - 0.017 mol
 - 0.25 mol
- What is the molality of a solution containing 0.5 moles of solute in 500 g of water?
 - 1 M
 - 1 m**
 - 0.001 M
 - 0.001 m
- A 40% alcohol solution contains...
 - 40 mL of alcohol in 100 mL of water
 - 100 mL of alcohol in 40 mL of water
 - 60 mL of alcohol in 100 mL of water
 - 40 mL of alcohol in 60 mL of water**
- How much 12 M HCl is needed to prepare 500 mL of 0.5 M solution of HCl?
 - 20.8 mL**
 - 229.2 mL
 - 250 mL
 - 270.8 mL
- Which of the following is a colligative property?
 - heating of a solvent
 - allowing a carbonated beverage to warm to room temperature
 - adding to salt into the water that pasta is being cooked in**
 - pouring a concentrated solution into a dilute solution
- Water is known as...
 - the oddity of molecules
 - the universal solvent**
 - the great dissolver
 - the molecule of living things
- Which of the following molecules are polar?
 - CF_4
 - CO_2
 - PCl_3**
 - PH_3
- Which of the following compounds is most likely to dissolve in water?
 - CH_4
 - $CaCl_2$**
 - I_2
 - $C_{12}H_{22}O_{11}$

